

# Chemical Reduction of Titania Supported Gold by Glycerol

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**Abstract:** Gold was loaded on TiO<sub>2</sub> by deposition precipitation method (DPU). The as prepared DPU samples were reduced in the liquid phase through chemical reduction with pure glycerol. The advantage of this method is that, the formation of small particles (less than 7 nm) can be prevented. The chemically reduced 0.93% Au/TiO<sub>2</sub> sample exhibited an average gold size of 12.6 ± 0.92 nm while, the H<sub>2</sub>-reduced counterpart showed 4.2 ± 0.25 nm. The efficacy of the size-controlled catalyst was demonstrated in the selective oxidation of HMF. The chemically reduced samples exhibited excellent selective oxidation performance by forming exclusively one product. This study reveals the possibility of controlling the size of metal-oxide supported gold catalysts.

**Keywords:** Gold nanoparticles, Titania, chemical reduction, glycerol, selective oxidation, size-controlled synthesis.

## I. INTRODUCTION

Titanium dioxide (TiO<sub>2</sub>) is considered as one of the most interesting transition metal oxide support in heterogeneous catalysis due to its reactivity, eco-friendliness, low cost and abundance [1]. Loading TiO<sub>2</sub> with gold nanoparticles leads to many exciting catalytic applications including CO oxidation in the gas phase [2-4] and photocatalytic activity in the liquid phase [5-8]. Due to the localized surface plasmon resonance (LSPR) the gold nanoparticles help titania to strongly absorb visible light. But, the LSPR properties are influenced strongly by the characteristics of gold nanoparticles such as shape [9], size [10-12] and gold loading [13]. Clearly, these characteristics of gold nanoparticles can be used as tools to tailor the catalytic properties of Au/TiO<sub>2</sub> catalysts.

However, the gold size requirement depends on the type of reaction. Supported gold catalysts with small gold size (2–3 nm) exhibited the best CO oxidation activity [2-4]. But small gold particles are not required for all reactions. Particularly, for the base-assisted selective oxidation reactions, medium sized particles (~8-10 nm) have shown better performance than that of smaller gold particles [14-15]. But in general, the net catalytic activity is obtained as the sum of dispersed atoms, small clusters, nanoparticles of varying sizes. This is the main problem in the activity-size correlation. It would be better to have narrow size distribution and the monodispersity would be the best case which remains only imaginable. In order to control the size of metal oxide-supported gold nanoparticles, two options have been attempted. One is size control through gold loading [12, 13].

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But this can work up to certain limits only and again the histogram looks with broader size distributions, just in different size ranges. In general, the Au(III)/TiO<sub>2</sub> samples are either calcined or reduced. Note that, thermal reduction under flowing hydrogen always generates only small particles ( $\leq 5$  nm) even with high gold loadings such as 5% Au/TiO<sub>2</sub>. Therefore, the current challenge is to control the size/ shape of oxide-supported gold nanoparticles. Here, we report the preparation of Au/TiO<sub>2</sub> catalyst without the presence of small gold particles achieved through the techniques of chemical reduction method. In this method, glycerol is used as reducing agent. No other chemicals are needed. The advantages of glycerol include non-toxicity, non-hazardousness, non-volatility, biodegradability that is generated as a byproduct in the manufacture of biodiesel fuel from vegetable oils. One of the attractive physical properties of glycerol is its high boiling point (290 °C). Furthermore, its high viscosity (1.412 Pa.s) [16] can prevent fine particles from settling down. This study demonstrates that, size-controlled synthesis of oxide-supported gold catalysts is possible by chemical reduction method. The results show that, even with low loading like 1 wt.% Au, it is possible to prevent the formation of nanoparticles below 7 nm, which is very interesting for selective oxidations.

## II. EXPERIMENTAL SECTION

### A. Chemicals and Techniques

HAuCl<sub>4</sub>.4H<sub>2</sub>O (SHOWA, 99.9%), glycerol (Aldrich, 99%), TiO<sub>2</sub> (Degussa, 50 m<sup>2</sup>g<sup>-1</sup>), and urea (Aldrich, 98%) were used as received. The glassware used in the synthesis was cleaned with aqua-regia and then rinsed with deionized water and dried overnight prior to use. Chemical analyses were performed by duo inductively coupled plasma atom emission spectroscopy (ICPAES) with Thermo Scientific iCAP 6500 instrument. The Au loadings were expressed as wt%. The TEM study done with a TECNAI G2 microscope. UV-Vis analysis were performed out with SHIMADZU UV-1800 equipment.

### B. Preparation of Catalyst

TiO<sub>2</sub> was dispersed in distilled water and the slurry temperature was kept 80 °C. Solution of HAuCl<sub>4</sub>.4H<sub>2</sub>O was added. It was followed by urea addition. The mixture was heated at 80 °C/12h. Centrifugation was performed to recover the solid, washed 3 times with water and dried in a desiccators. The aliquots of dried (*as prepared*) samples were treated in glycerol at different temperatures to obtain reduced gold catalyst. In a typical process, 150 mg of finely powdered *as prepared* samples was and mixed with 30 g of pure glycerol in a closed 50 mL vial and stirred vigorously.

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The catalyst mixed well with the glycerol due to the presence of residual moisture in the *as prepared* catalyst and hygroscopic nature of glycerol. The reaction mixture was stirred at 100 °C in an oil bath for 12 h. The progressive reduction of gold could be followed with changes in color (which is not possible with conventional gas-phase catalyst activation). After stirring for 12 h, the reaction mixture was removed from oil bath, cooled and diluted with excess of water. Centrifugation was performed to recover the catalysts, followed by washing (3 times with deionized water) and finally mixed with the reaction mixture immediately. For the purpose of comparison, thermal reduction under flowing H<sub>2</sub> was performed (at 300 °C/2h) with the *as prepared* catalysts.

### C. Chemicals and Techniques

Aerobic oxidation of 5-Hydroxymethyl-2-furfural (HMF) was performed with supported gold catalysts. In a typical reaction, 0.1g HMF was dissolved in 100 mL distilled water and transferred into the stainless steel reactor followed by the addition of base (0.125 mg) and catalyst (150 mg). The reactor was closed with screw cap very tightly and 40 bar air pressure was applied. After stirring at 25 °C for 1h, the reaction was stopped. The product analysis was done by HPLC [Alltech OA-1000 column (300 mm × 6.5 mm) with eluent 0.1% H<sub>3</sub>PO<sub>4</sub>. 0.5 mL reaction mixture was diluted to a total of 5.0 mL with eluent. by UV (210 nm) and refractive index detectors used to identify products.

### III. RESULTS AND DISCUSSION

The chemical analyses performed by ICPAES shows that, the gold loading of the catalysts is 0.93 wt% Au/TiO<sub>2</sub> which is close to the nominal loading of 1%. Fig. 1 shows the TEM image and gold size distribution of 0.93%Au/TiO<sub>2</sub> sample reduced by chemical reduction with glycerol (CRG) at 100 °C/12h.

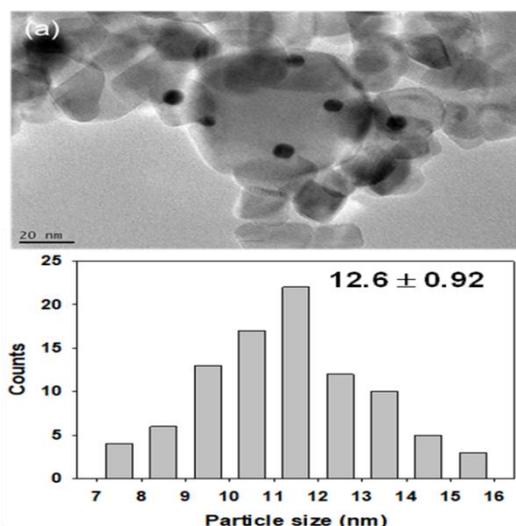


Fig. 1. TEM image and gold size distribution chemically reduced 0.93Au/TiO<sub>2</sub> catalyst.

Note that, the lowest gold size obtained lies above 7 nm, which a remarkable achievement in the size-controlled synthesis of oxide-supported gold nanoparticles. The 0.93 wt% Au/TiO<sub>2</sub> exhibits an average size of 12 ± 0.96 nm. Very large particles above 17 nm are also clearly absent. Interestingly, the gold loading is not very high (0.93 wt.%)

only) but, the average size of gold is above 10 nm, which is the unique advantage of this chemical reduction method.

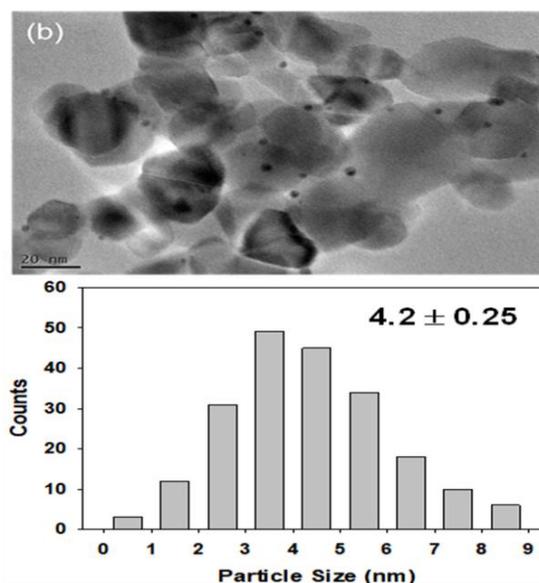


Fig. 2. TEM image and gold size distribution of 0.93Au/TiO<sub>2</sub> catalyst, reduced by flowing H<sub>2</sub> at 300 °C/2h.

That is, not only larger gold particles can be grown, but also, smaller particles can be effectively avoided. The same sample reduced under flowing H<sub>2</sub> (at 300 °C/2h) results into an average size of 4.2 ± 0.25 nm, which is normal as expected. Compare the Fig. 1 with Fig. 2. The TEM results reveal that, the characteristics of gold particles vary greatly with the applied reduction method.

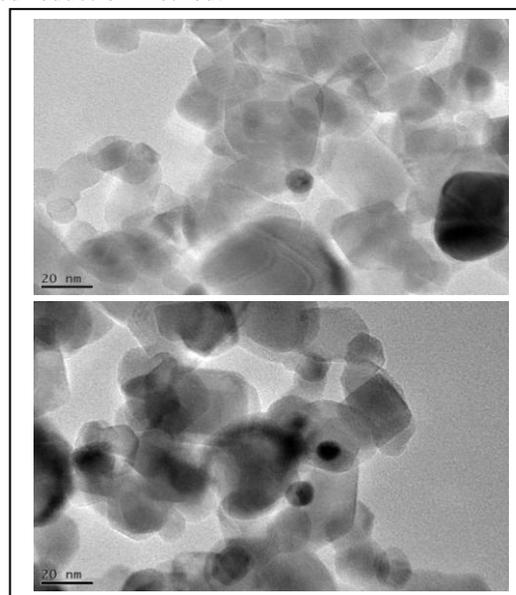


Fig. 3. TEM images of chemically reduced 0.93Au/TiO<sub>2</sub> catalyst.

As could be seen from the Fig.-1, the sample reduced by CRG method produced larger gold particles than that of H<sub>2</sub>-reduced samples. Note that, the gold size distribution is entirely in different ranges for these samples.

TEM images of chemically reduced sample were carefully examined and it was found that, the smaller gold particles below 7 nm were clearly absent (Fig. 3). These results indicate that, the growth of large particles with low gold loading is possible with CRG method, which is not possible with the conventional method of reducing under flowing H<sub>2</sub>. Furthermore, the present study reveals the efficacy of glycerol as a reducing agent.

It may appear quite different approach. When the as prepared DPU sample was mixed with glycerol, it was observed that, the fine powders mixed very well with glycerol by stirring [Fig. 4]. The driving force for such an observation could be related to the hygroscopic nature of glycerol and the residual moisture present in the as prepared DPU samples. In order to understand the growth of gold nanoparticles during chemical reduction method, sample drawn in between beginning and finished catalyst was examined [Fig. 5]. As could be seen from the figure, first small clusters appear and later, they merge to form larger particles, which is in line with the absence of smaller gold particles.

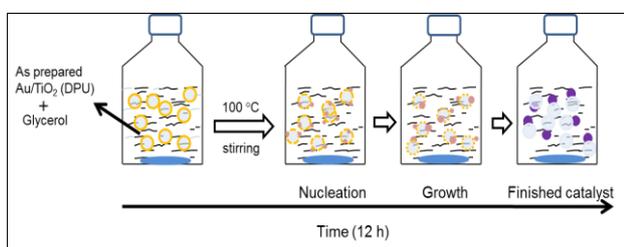


Fig. 4. Chemical reduction of gold over titania.

In fact, glycerol has been used to reduce various metals (but unsupported metals). Glycerol has been used as reducing agent for the ions of silver, gold and nickel. Gold nanoparticles have been obtained by reduction of HAuCl<sub>4</sub> encapsulated with liposome [17]. Gold was reduced in glycerol under both microwave and reflux conditions [18]. Nickel nanoparticles have been prepared in glycerol at 100 °C by using hydrazine hydrate in alkaline medium [19]. Silver ions in silver alkoxide has been reduced by glycerol to form silver nano-particles (D av. < 63 nm) at high temperature [20]. Nalawade et al [21] applied pure glycerol to reduce HAuCl<sub>4</sub> at room temperature under basic conditions (HAuCl<sub>4</sub> : NaOH mole ratio = 1:5). Gold size varied from 8 to 50 nm [21]. Interestingly, the concentration of glycerol played key role in determining the characteristics of resulting nanoparticles. 20:80 (vol/vol) mixture of glycerol and water produced the best results in the absence of stabilizers like PVP [21]. The advantage of the present method is that, it does not require any other reagent than pure glycerol, becoming a green protocol. It should be mentioned here that, in our earlier study, this technique worked at even room temperature with nano ceria support. However, for the Au/TiO<sub>2</sub> samples, it did not work at room temperature.

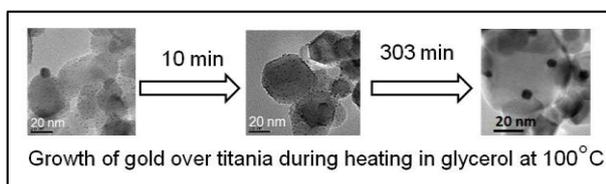
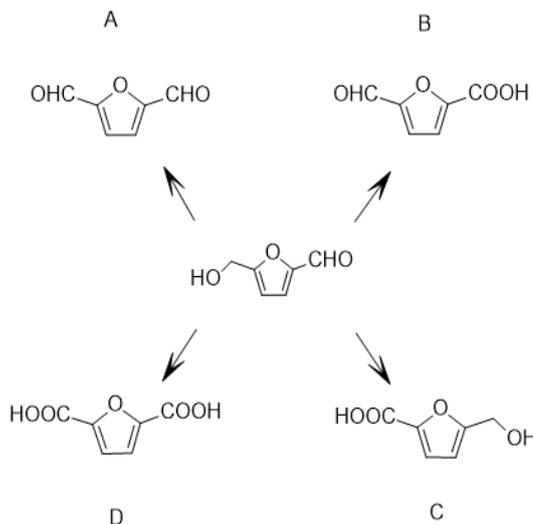


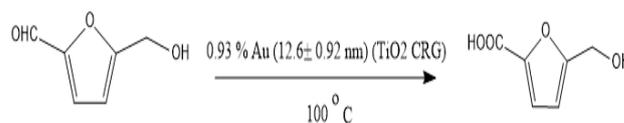
Fig. 5. Time dependant TEM analysis of 0.93Au/TiO<sub>2</sub> catalyst during chemical reduction process.

The striking advantage of this chemical reduction method was revealed in the catalytic activity tests. Among various supports, gold supported on TiO<sub>2</sub> was proved to be effective for HMF oxidation into 2,5-Furandicarboxylic acid (FDCA) [22]. According to Casanova *et al* [22], HMF oxidation to 5-hydroxymethyl-2-furancarboxylic acid (HMFCFA) is a very fast reaction.



Scheme. 1: Various oxidation products possible from HMF

The conversion of HMF and the product distribution depends strongly on the reaction conditions. The different types of oxidation products from HMF are depicted in the Scheme-1. According to Casanova *et al* [22] and Pasini *et al* [23], initially HMF is rapidly converted into HMFCFA and it is a primary and unstable product that is gradually converted into FDCA, which appeared as a secondary and stable product. It has been pointed out that, FDCA was practically not formed before all HMF was converted into HMFCFA, indicating the slower reactivity of HMFCFA compared with HMF. Higher reaction temperature and higher base concentration boost the undesirable rapid degradation of HMF into by products [24]. On the other hand, at lower temperatures, HMFCFA was the major product at complete conversion of HMF (Scheme-2).

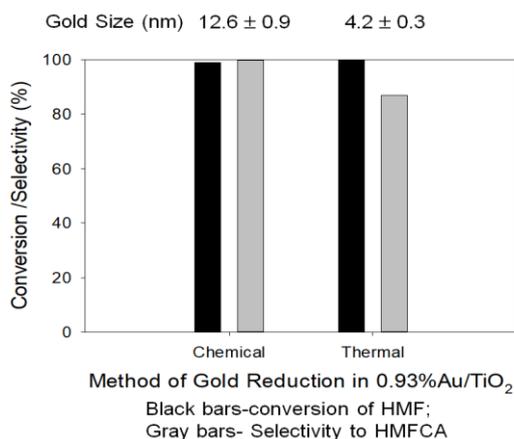


Scheme. 2 : Selective oxidation of HMF over size-selective Au/TiO<sub>2</sub> catalyst.

At 60 °C, 1% Au (5.2 nm)/TiO<sub>2</sub> catalyst produced 20% and 80% yields of FDCA and HMFCFA (Reaction conditions: 240 min, 10 bar of oxygen pressure, HMF : metal loading : NaOH molar ratio 1 : 0.01 : 4) respectively [23]. The yield of FDCA was >99% with the Au/TiO<sub>2</sub> in an another study reported by Casanova *et al* [22] (Reaction conditions:

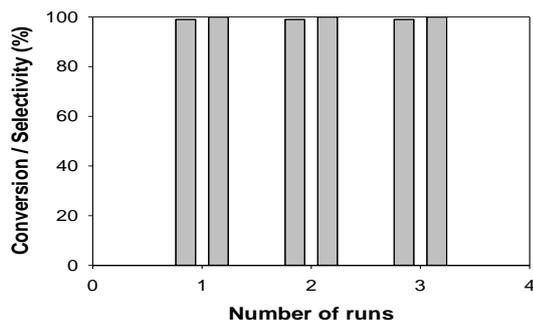
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HMF (3 mmol) + NaOH (12 mmol) in 20 mL H<sub>2</sub>O, the reaction mixture was loaded into an autoclave bubbling air flow 0.5 mL s<sup>-1</sup> at 10 bar). 1.6% Au (2.6 nm)/TiO<sub>2</sub> [25] exhibited full conversion of HMF at 22 °C with 8% and 92% selectivities to FDCA and HMFCFA respectively (0.15M HMF solution in 0.3M NaOH; metal : HMF = 6.67×10<sup>-3</sup> mol/mol, T = 22 °C, P = 690 kPa O<sub>2</sub>. Reaction time = 6 h).



**Fig. 6. Conversion and selectivity of differently reduced gold catalysts towards oxidation of HMF.**

Selective oxidation reaction of HMF was studied with chemically reduced samples and compared with that of H<sub>2</sub>-reduced samples. Figure-6 shows the comparative results of selective oxidation of HMF under mild conditions. The 0.93% Au/TiO<sub>2</sub> sample reduced by H<sub>2</sub> shows that, at full HMF conversion, the main product is 5-hydroxymethyl-2-furfural (HMFCFA) along with FDCA as a minor product. This result is in accordance with the previous reports [23, 25]. On the other hand, the chemically reduced counterpart showed that, HMFCFA was the only product and the FDCA formation is suppressed to trace (at 99% HMF conversion). This can be attributed to the relatively larger size of gold particles in the case of chemically reduced samples which catalyzed very mild oxidation. The suppressing of FDCA formation is in accordance with the absence of small gold particles (<7nm) (Fig. 3) and in agreement with the aforementioned literature reports [14, 15, 26, 27, 23]. In order to ascertain the catalytic performance of 0.93% Au/TiO<sub>2</sub> sample reduced by CRG method, recycling tests were performed. The used sample was washed with aq. NaOH (50 mL of 0.3125M NaOH) once and distilled water (3 times) and immediately transferred into the fresh reaction mixture.



**Fig. 7. Results of reproducibility tests of HMF selective oxidation with 0.93% Au(12.6±0.9 nm)/TiO<sub>2</sub> catalyst.**

Figure 7 shows the results obtained from the re-cycling tests. The conversion of HMF did not decrease and the product distribution is the same, showing the effectiveness of the chemical reduction method to prepare Au/TiO<sub>2</sub> catalysts. A catalytic run was conducted in the absence of the catalyst. Without catalyst, only very low conversion of HMF occurred and only the Cannizzaro products were obtained.

## IV. CONCLUSIONS

The process of chemical reduction with glycerol has been demonstrated with the *as prepared* Au/TiO<sub>2</sub> (DPU) samples. Glycerol effectively reduces titania supported gold phase at 100 °C. Smaller gold particles could be successfully prevented, which is a remarkable advantage for the selective oxidation catalysis. Compared with the conventional thermal reduction under flowing hydrogen at 300 °C, relatively larger gold nanoparticles can be prepared by CRG method at 100 °C. The catalytic efficiency of chemically reduced Au/TiO<sub>2</sub> samples has been examined in the selective oxidation of 5-hydroxymethyl-2-furfural under very mild conditions. The catalytic results revealed the advantage that, Au/TiO<sub>2</sub> samples (CRG) exhibited very mild and controlled oxidation capacity than that of Au/TiO<sub>2</sub> (H<sub>2</sub>/300 °C) due to larger size. Only product obtained was HMFCFA. This study throws light on the size-controlled synthesis of metal oxide-supported gold catalysts which is crucial for selective oxidations.

## V. ACKNOWLEDGEMENTS

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## REFERENCES

1. X. Chen and S. Mao, "S. Titanium Dioxide Nanomaterials: Synthesis, Properties, Modifications, and Applications". 2007. *Chem. Rev.* 107, pp 2891–2959.
2. F. Cosandey and T. E. Madey, 2001 *Surf. Rev. Lett.*, 8, pp 73–83).
3. G. R. Bamwenda, S. Tsubota, T. Nakamura and M. Haruta, 1997. *Catal. Lett.* 44, pp 83–88.
4. M. Valden, X. Lai and D. W. Goodman, 1998, *Science.*, 281, pp. 1647–1650.
5. Zhang, D. Q. *et al.* Au Nanoparticles Enhanced Rutile TiO<sub>2</sub> Nanorod Bundles with High Visible-Light Photocatalytic Performance for NO Oxidation. 2014, *Appl. Catal. B-Environ.* 147, pp 610–616.
6. Nalbandian, M. J. *et al.* Tailored Synthesis of Photoactive TiO<sub>2</sub> Nanofibers and Au/TiO<sub>2</sub> Nanofiber Composites: Structure and Reactivity Optimization for Water Treatment Applications. 2015, *Environ. Sci. Technol.* 49, pp 1654–1663.
7. M. McEntee, A. Stevanovic, W. J. Tang, M. Neurock, and J. T. Yates, Electric Field Changes on Au Nanoparticles on Semiconductor Supports—the Molecular Voltmeter and Other Methods to Observe Adsorbate-Induced Charge-Transfer Effects in Au/TiO<sub>2</sub> Nanocatalysts". 2015, *J. Am. Chem. Soc.* 137, pp1972–1982.
8. Tsukamoto, D. *et al.* "Gold Nanoparticles Located at the Interface of Anatase/Rutile TiO<sub>2</sub> Particles as Active Plasmonic Photocatalysts for Aerobic Oxidation". 2012, *J. Am. Chem. Soc.* 134, pp 6309–6315.
9. Ma, L. *et al.*, "Synthesis of Dumbbell-Like Gold-Metal Sulfide Core-Shell Nanorods with Largely Enhanced Transverse Plasmon Resonance in Visible Region and Efficiently Improved Photocatalytic Activity", 2015, *Adv. Funct. Mater.* 25, pp898–904.

10. Zielinska-Jurek, A. *et al.*, "Preparation and Characterization of Monometallic (Au) and Bimetallic (Ag/Au) Modified-Titania Photocatalysts Activated by Visible Light", 2011, *Appl. Catal. B-Environ.* 101, pp504–514.
11. K. Kimura, S. Naya, and Y. Jin-nouchi, TiO<sub>2</sub> "Crystal Form-Dependence of the Au/TiO<sub>2</sub> Plasmon Photocatalyst's Activity". 2012, *J. Phys. Chem. C* 116, pp 7111–7117.
12. M. Murdoch, G. I. N. Waterhouse and M. A. Nadeem, "The Effect of Gold Loading and Particle Size on Photocatalytic Hydrogen Production from Ethanol over Au/TiO<sub>2</sub> Nanoparticles", 2011, *Nat. Chem.* 3, pp 489–492.
13. A. Tanaka, S. Sakaguchi, and K. Hashimoto, "Preparation of Au/TiO<sub>2</sub> with Metal Cocatalysts Exhibiting Strong Surface Plasmon Resonance Effective for Photo-induced Hydrogen Formation under Irradiation of Visible Light", 2012, *ACS Catal.* 3, pp79–85.
14. N. Dimitratos, J. A. Lopez-Sanchez, D. Lennon, F. Porta, L. Prati and A. Villa, *Catal. Lett.*, 2006, 108, pp147–153.
15. P. Haider, B. Kimmerle, F. Krumeich, W. Kleist, J.D. Grunwaldt and A. Baiker, *Catal. Lett.*, 2008, 125, pp169–176
16. Segur, J. B.; Oberstar, H. E. 1951, Viscosity of Glycerol and Its Aqueous Solutions". *Industrial & Engineering Chemistry.* 43 (9): pp2117–2120.
17. R. Genc, G. Clergeaud, M. Ortiz and C. K. O'Sullivan, "Green Synthesis of Gold Nanoparticles Using Glycerol Incorporated Nanosized Liposomes," *Langmuir*, Vol. 27, No. 17, 2011, pp. 10894-10900.
18. N. Grace and K. Pandian, "One Pot Synthesis of Polymer Protected Gold Nanoparticles and Nanoprisms in Glycerol," 2006, *Colloids and Surfaces A: Physicochemical and Engineering Aspects*, 290 pp138-142.
19. K. Singh, K. H. Katiel, V. V. S. Chilukuri and P. K. Khanna, "Glycerol Mediated Low Temperature Synthesis of Nickel Nanoparticles by Solution Reduction Method, 2011, *J. Nanosci. Nanotechnol.*, Vol. 11, No. 6, , pp. 5131-5136.
20. E. Nisaratanaporn and K. Wongsuwan, "Preparation of Ultrafine Silver Powder Using Glycerol as Reducing Agent," 2008, *J. Metals, Materials and Minerals*, Vol. 18, No. 2, pp. 1-5.
21. Pradnya Nalawade, Tulsi Mukherjee, Sudhir Kapoor, 2013, *Advances in Nanoparticles*, 2, pp78-86.
22. S. Casanova, S. Iborra, A. Corma, *Chem Sus Chem.*, 2009, 2, pp1138–1144.
23. Pasini, M. Piccinini, M. Blosi, R. Bonelli, S. Albonetti, N. Dimitratos, J. A. Lopez- Sanchez, M. Sankar, Q. He, C.J. Kiely, G. J. Hutchings, F. Cavani. *Green Chem.*, 2011, 13, pp2091–2099.
24. C. A. Antonyraj, N. T. T. Huynh, S. K. Park, S. Shin, Y. J. Kim, S. Kim, K. Y. Lee, J. K. Cho, *Appl. Catal. A.*, 2017, 547, pp230–236.
25. S. E. Davis, L. R. Houk, E. C. Tamargo, A.K. Datye, R. J. Davis, 2011, *Catal. Today*. 160 pp55–60.
26. Y.R. Leshkov, C. J. Barret, Z.Y. Liu, J. A. Dumesic, *Nature.*, 2007, 447, pp 982–985.
27. A. Corma, S. Iborra, A. Velty, *Chem. Rev.*, 2007, 107, pp2411–2502.

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**Chennan Ramalingan** profile which contains their education details, their publications, research work, membership, achievements, with photo that will be maximum 200-400 words.