

# Design Of Galvanic Cell Battery For Underwater Applications Using Seawater As Electrolyte

S.A.Thirumalini, D.Deepika, K.Muthumeenakshi, S.SakthivelMurugan

**Abstract:** Seawater battery is one of the green electricity sources to fulfill energy need for electrical equipment, especially in the coastal area and fishing activity. A survey was conducted among fishermen in which it was found that small scale fishermen uses lead acid battery and fuel cell in order to charge the mobile phone and glow the fishing lights. But major drawback of lead acid battery and fuel cell is that the maintenance cost is higher which is difficult for them to afford. Seawater is one most available sources all over the world and it is of no cost, hence a seawater battery is designed. This paper aims to study galvanic cells using sea water as electrolyte for energy harvesting. The electrochemical performances of Galvanic cells were carried out by measuring electric potentials by understanding the nature of conductivity of electrodes. The effect of sea water pH on electric potential was analyzed using sea water from different parts of Bay of Bengal with varying depths. Various combinations of electrodes like Graphite, Zinc, Copper, Aluminium, Brass and Iron were tested. A maximum yield of 1.1 V was obtained using the combination of Graphite-Iron as Cathode-Anode for a single cell. Further, we developed a working prototype for 16 cell. It generates a voltage of 12 V and 20 mA. Since the output current obtained was not as desired so we added a current amplification circuit and obtained a maximum current of 300mA from 20mA.

**Keyword:** Galvanic cell, seawater, Current amplification, electrolyte.

## I. INTRODUCTION

Galvanic cell is an electrochemical device which directly converts chemical energy into electrical energy as long as electrolysis process exist. Galvanic cells have attracted much attention as an alternative energy source for portable electronic, electrical batteries. The secondary cell and fuel cell are mainly developed to avoid the disadvantages of generator. But, these fuel cell and secondary cell has various disadvantages like oil spill ,high cost of maintenance etc. In order to overcome these disadvantages, a galvanic cell battery is developed [1][2]

This project is to develop a stack arrangement of galvanic cell battery to provide the power, in order to charge the mobile phone and to glow the fishing light which are two major factor necessary for small scale fishermen day to day life.

Here the electrodes are selected and tested to find out which combination of electrode gives us good potential. Here sea water is used as electrolyte[3]. After the selection of electrode ,then it is tested to find which arrangement and connection help us to achieve good potential. Then it subjected under various condition. Since the current

produced by the battery is very less, current boosting circuit have been developed and validated in order to achieve the desired potential.

## II.METHODLOGY AND WORKING

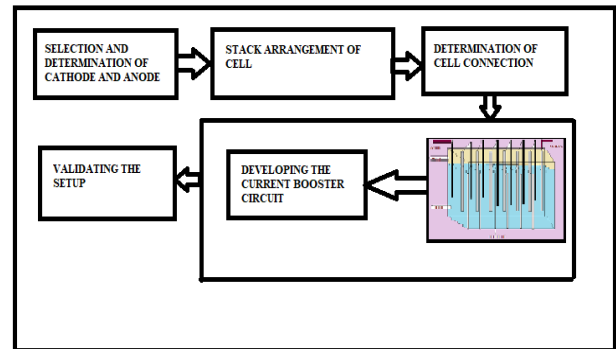


Fig 1: Design of Galvanic cell Battery

The Fig 1 shows the design of galvanic cell battery. At first the anode and cathode are selected which produces maximum potential. Then cells are arranged in stack as well provided with series connection Finally booster circuit are added in order to achieve the desired voltage

### Analyzing the Various Electrode

The choice of electrode like copper, zinc, aluminum and graphite is analyzed using seawater as electrolyte[4].It has been tested with different sea water sample from different location like Pondicherry, Cuddalore, Mahabalipuram and Chennai which is collected at different depth like 5m, 10m,20m .It generate an voltage of about 1V.Electrode were tested in different combination and found that Graphite-Iron, Graphite-Zinc produces good potential. Fig 2 shows that the performance of electrode in sea water as electrolyte. Here, carbon is used as cathode and tested with variable anode. Finally it is concluded with Graphite-Iron combination. Moreover rusting Graphite take place when zinc is used as anode.

The Table II shows the quantitative analysis of cell. The obtained voltage from TABLE II is based on number of cells connected in series combination. From this it is understand that as number cell increases , voltage also increases.

TABLE III Performance Analysis of various cell under various condition

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CONDITION OF THE CELL	TIME PERIOD OF EXPOSURE (HOURS)	VOLTAGE OBSERVED (V)
Without Zinc metal	23	12.03
With Zinc metal	288	10.63

Table III shows the performance of 16 cell battery with and without addition of poison material. It is found that using the Zinc as Poison material it withstand the potential of about 10 V for about 288 hours.

### Calculation of corrosion of iron

The rate of corrosion is the speed at which any given metal deteriorates in a specific environment. The rate, or speed, is dependent upon environmental conditions as well as the type of the metal [13].

Corrosion rate is given by equation (5)

$$\text{Corrosion Rate} = 87.6 \times (W / (\text{DAT})) \quad (5)$$

where,

W= Weight loss in milligram

D = Metal density in g/cm<sup>3</sup>

A = Area of sample in cm<sup>2</sup>

T = Time of exposure of metal sample in hours

### CURRENT BOOSTER'S RESULTS

As obtained current is insufficient it is necessary to amplify the current using current boosting circuits. This can be achieved using multisim software. The internal resistance calculated for the above battery is 600 Kilo Ohm as shown in below Fig 3.3 The amp hour is calculated under idealistic condition .it is found that it has a capacity of 0.46Ah . It is noted that battery will produce constant current of 20 milliamperes for 23 hours.By using the above parameters, it is found that Current(A):20mA,Constant time:23hr

#### Amp-Hr Calculation

$$\text{Amp-Hr (Ah)} = \text{Current(A)} * \text{Time(Hr)}$$

$$= 0.02 * 23$$

$$= 0.46 \text{Ah or } 460 \text{mAh}$$

After calculation of amp hour and internal resistance, these values are incorporated to design an non ideal battery. Various circuits are simulated using non ideal battery.It is found the Fig 4 show s the current boosting which produces desired current and voltage

#### Current Booster circuit using LM7812 and 1N4002

This Fig 4 shows the current booster circuit using LM7812 and 1N4002. In general,LM7812 not only regulate the voltage,but also amplify the current.1N4002 amplify the current under forward bias condition.LM7812 regulator along with 1N4002 diode amplify the current to about 1A. It produces an output voltage 12.54V and 1.12A.

The output is validated in three phases

### Validation of bottom module

Battery model is developed using acrylic material. The major advantage of using acrylic material container is that, it reduces water leakage when compared to other material containers. The prototype model of 16 cells is shown in the Fig 5 below

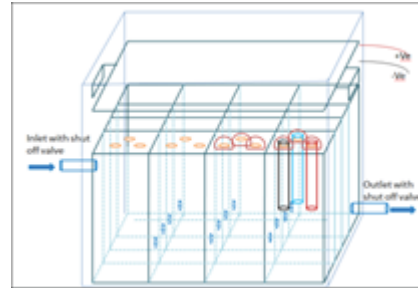


Fig 5: Prototype model of 16 CELLS

The Fig 5 shows the prototype model of 16 cell battery.

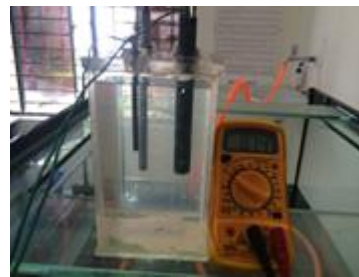


Fig 6: Customized single cell setup

The Fig 6 shows the Customized model of single cell battery. Further this setup has been extended to 16 cell as shown in Fig 7



Fig 7: Customized single cell setup

Finally the setup have executed for 16 cell as shown in Fig 7.It produce an output voltage of 12V and current of 20 milliamperes

### Validation of top module

Further the circuit are tested using a ordinary battery two 9

V battery. The input current produced by this battery is similar to that of battery that we developed and is approximately 20 milliamperes

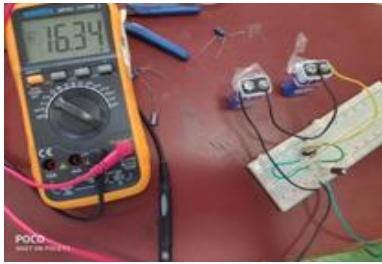


Fig 8: Circuit using LM7812 in an ordinary battery

The Fig 8 shows the circuit using ordinary battery LM7812. It is found that the voltage obtained from the battery is 16.3 volt and current obtained from the battery is 18.2 milliamperes which is comparatively very less. Hence circuits need further improvement.

#### Validation of top and bottom module

The current booster circuit has been designed after making certain modification in the circuit like replacing 1N4002 with 330 kilo ohm resistor and replacing the load with 10 ohm. It is found that for 16 cell it produces an voltage of about 11.92 V and current of 300 milli ampere.



Fig 9: Modified Current booster Circuit incorporating top and bottom module

Fig 9 shows the voltage obtained after incorporating both top and bottom module. It is found that it produces an voltage of about 11.92 V. The current obtained after incorporating both top and bottom module. It is found that it produces an voltage of about 12V and current of about 300 milliamperes. Though, various modifications are made in the current booster circuit, we can only achieve only 300 milliamperes and at the same time voltage drops to 11.92V. Both the values are inefficient, since it cannot charge the mobile phone or glow the fishing lights. Therefore further improvement is needed for the circuit.

#### IV. CONCLUSION

This paper was started with motivation to provide a sufficient power mobile phone charger and to glow the fishing lights. The approximate voltage and current required is 12V and 1A. With seawater as electrolyte, it is concluded that GRAPHITE-IRON combination produces good potential.

The customized battery of 16 cells produces a voltage of 12V and a current of only 20 milliamperes, which is very less. Therefore current boosting circuits are used to obtain a higher value of current. After addition of a current booster circuit, the battery produces the voltage of 12V and a current of about 300 milliamperes.

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## DESIGN OF GALVANIC CELL BATTERY FOR UNDERWATER APPLICATIONS USING SEAWATER AS ELECTROLYTE

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